TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Plumbi Carbonas,	Calcium (chalk).	Oxalate of Ammonium after removing Lead.
	Sulphate of Barium or Lead.	Insolubility in Acetic Acid.
	Silicates.	Insolubility in Acetic Acid.
	Alkaline Salts.	Evaporation after removing Lead.
Plumbi Iodidum,	Chromate (of Lead).	Solubility in Chloride of Ammonium.
	Zinc, etc.	Evaporation after removing Lead.
Plumbi Nitras,	Zinc, etc.	Evaporation after removing Lead.
	Copper.	Excess of Ammonia.
Plumbi Oxidum,	Carbonate.	Dilute Acids.
	Zinc, etc.	Evaporation after removing Lead.
Potassa,	Organic matter.	Color of solution, and Permanganate of Potash.
	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
	Carbonate.	Effervescence with acids.
	Silica.	Solubility in Alcohol.
Potassa cum Calce,	Silica.	Solubility in Hydrochloric Acid.
Potassa Sulphurata,	Deficiency of Sulphide.	Sulphuretted Hydrogen.
Potassii Acetas,	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
	Silica.	Evap'n of acid solution, insolubility of residue.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
	Alkaline earths.	Carbonate of Sodium.
	Carbonate.	Effervescence with Acetic Acid.
	Organic impurities.	Sulphuric Acid.
Potassii Bicarbonas,	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Carbonate.	Chloride of Barium in the cold.
Potassii Bichromas,	Sulphate.	Chloride of Barium.
Potassii Bitartras,	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
	More than 6 per cent. of Tartrate of Calcium.	Quantitative Analysis. " "
Potassii Bromidum,	Bromate.	Sulphuric Acid.
	Iodide.	Chlorine and mucilage of Starch.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Potassii Bromidum,	Sulphate.	Chloride of Barium.
	More than 3 per cent. of Chloride.	Quantitative Analysis.
Potassii Carbonas,	Silica, etc.	Insolubility of residue on evaporation of acid solution.
	Alkaline earths.	Carbonate of Sodium.
	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
Potassii Chloras,	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Calcium.	Oxalate of Ammonium.
Potassii Citras,	Carbonate.	Effervescence with acids.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Tartrate.	Acetic Acid.
Potassii Cyanidum,	Carbonate.	Effervescence with Acids.
Potassii et Sodii Tartras,	Calcium.	Oxalate of Ammonium.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
Potassii Ferrocyanidum,	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
Potassii Hypophosphis,	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Phosphate.	Magnesia mixture.
	Calcium.	Oxalate of Ammonium.
Potassii Iodidum,	Iodate.	Mucilage of Starch and Tartaric Acid.
	Chloride or Bromide.	Ammonia, Nitrate of Silver, and Nitric Acid.
	Sulphate.	Chloride of Barium.
Potassii Nitras,	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
	Alkaline earths.	Carbonate of Ammonium.
	Sulphate.	Chloride of Barium.
	Chloride (of Sodium).	Nitrate of Silver.
Potassii Permanganas,	Nitrate.	Sulphuric Acid and Ferrous Sulphate to decolorized solution.
	Chloride.	Nitrate of Silver to decolorized solution.
	Sulphate.	Nitrate of Barium after removing Manganese by Ammonia.
Potassii Sulphas,	Alkaline earths.	Carbonate or Phosphate of Ammonium.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
	Chloride.	Nitrate of Silver.
Potassii Sulphis,	Sulphate.	Chloride of Barium.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Potassii Tartras,	Calcium.	Oxalate of Ammonium.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
Quinidinæ Sulphas,	Organic impurities.	Sulphuric Acid.
	Morphine.	Nitric Acid.
	Cinchonine, Quinine, or Cinchonidine.	Iodide of Potassium and Ammonic Hydrate.
Quinina,	Organic impurities.	Sulphuric Acid.
	Cinchonine, Cinchonidine, or Quinidine.	Sulphate of Ammonium and Ammonia.
Quininæ Bisulphas,	Organic impurities.	Sulphuric Acid.
	Free water.	Drying upon water-bath.
	Sulphates of Quinidine, Cinchonidine, or Cinchonine.	Ammonia, as for Quinine.
Quininæ Hydrobromas,	Organic impurities.	Sulphuric Acid.
	Free water.	Drying upon water-bath.
	Sulphate.	Chloride of Barium.
	Barium.	Sulphuric Acid.
	Cinchonine, etc.	Ammonia, as for Quinine.
Quininæ Hydrochloras,	Organic impurities.	Sulphuric Acid.
	Barium.	Sulphuric Acid.
	Sulphate.	Chloride of Barium.
Quininæ Sulphas,	Organic impurities.	Sulphuric Acid.
	Ammonia (Sulphate).	Boiling with milk of lime.
	Free water.	Drying on water-bath.
	Cinchonine Sulphate, etc.	Ammonium Hydrate.
Quininæ Valerianas,	Organic impurities.	Sulphuric Acid.
	Sulphate.	Chloride of Barium.
Rheum,	Turmeric.	Boracic Acid.
Saccharum,	Insoluble salts, etc.	Aqueous or alcoholic solution on standing.
	Grape-Sugar or Inverted Sugar.	Nitrate of Silver and Ammonic Hydrate.
Saccharum Lactis,	Cane-Sugar.	Sulphuric Acid.
Salicinum,	Mineral matter.	Incineration.
Santoninum,	Mineral matter.	Incineration.
Sapo,	More than 34 per cent. of water.	Drying at 110° C.
	Animal Fats.	Gelatinization of 4 per cent. alcoholic solution.
	Carbonate of Sodium.	Solubility in alcohol.
	Silica and insoluble matter.	Solubility in water.
	Metals.	Sulphuretted Hydrogen.
Sapo Viridis,	More than 4 per cent. of water.	Drying at 100° C.
	Free Fats.	Digestion of dried soap in Benzol.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Sapo Viridis,	Insoluble Carbonates.	Dilute Acids to residue from alcohol and water.
	Starch.	Iodine to residue from alcohol and water.
Scammonium,	Chalk.	Effervescence with Acids.
	Starch.	Iodine.
Scammonii Resina,	Resin of Guaiacum.	Inner surface of potato-paring.
	Resin of Jalap.	Insolubility in Ether.
Soda,	Organic matter.	Color of aqueous solution; Sulphuric Acid.
	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
	Carbonate.	Effervescence with Acids.
	Silica or Carbonate.	Solubility in Alcohol.
Sodii Acetas,	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
	Silica.	Insolubility of residue on evaporating acid solution.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
	Alkaline earths.	Carbonate of Sodium.
	Carbonate.	Effervescence with Acids.
	Organic impurities.	Sulphuric Acid.
Sodii Arsenias,	Arsenite.	Sulphuretted Hydrogen water.
	Excess or deficiency of water of crystallization.	Quantitative Analysis.
Sodii Benzoas,	(*Vide* Acidum Benzoicum.)	
Sodii Bicarbonas,	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
	Ammonium Salts.	Boiling with solution of Soda.
	Carbonate.	Chloride of Barium in the cold, and Quantitative Analysis.
Sodii Bisulphis,	Sulphate.	Chloride of Barium.
Sodii Boras,	Carbonate.	Effervescence with Acids.
	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
	Alkaline earths.	Carbonate of Sodium.
	Metals.	Hydrosulphuric Acid.
Sodii Bromidum,	Bromate.	Sulphuric Acid.
	Iodide.	Chlorine-water and mucilage of Starch.
	Sulphate.	Chloride of Barium.
	More than 3 per cent. of Chloride.	Quantitative Analysis.
Sodii Carbonas,	Chloride.	Nitrate of Silver.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Sodii Carbonas.	Sulphate.	Chloride of Barium.
	Metals.	Hydrosulphuric Acid.
	Alumina.	Ammonia and Chloride of Ammonium.
Sodii Chloras,	Potassium.	Bitartrate of Sodium.
	Calcium.	Oxalate of Ammonium.
	Chloride.	Nitrate of Silver.
	Sulphate.	Chloride of Barium.
Sodii Chloridum,	Metals.	Hydrosulphuric Acid or Sulphide of Ammonium.
	Alkaline earths.	Carbonate of Sodium.
	Sulphate.	Chloride of Barium.
	Iodide or Bromide.	Chlorine-water and Starch to residue on evaporating alcoholic solution.
Sodii Hypophosphis,	Calcium.	Oxalate of Ammonium.
	Potassium.	Bitartrate of Sodium.
	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Phosphate.	Ammonia and Sulphate of Magnesium.
Sodii Hyposulphis,	Sulphate.	Chloride of Barium.
	Carbonate.	Effervescence with Acids.
Sodii Iodidum,	Iodate.	Mucilage of Starch and Tartaric Acid.
	Sulphate.	Chloride of Barium.
	Chloride or Bromide.	Ammonia, Nitrate of Silver, and Nitric Acid.
Sodii Nitras,	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
	Alkaline earths.	Carbonate of Ammonium.
	Potassium.	Bitartrate of Sodium.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Iodide.	Chlorine-water and Mucilage of Starch.
Sodii Phosphas,	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
Sodii Pyrophosphas,	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
Sodii Salicylas,	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Organic impurities.	Sulphuric Acid.
Sodii Santoninas,	Alkaline earths.	Carbonate of Sodium.
	Alkaloids.	Precipitate with Tannic or Picric Acid.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Sodii Sulphas,	Carbonate.	Effervescence with Acids.
	Chloride.	Nitrate of Silver.
	Metals.	Hydrosulphuric Acid or Sulphide of Ammonium.
	Ammonium Sulphate.	Boiling with Soda.
Sodii Sulphis,	Sulphate.	Chloride of Barium.
Sodii Sulphocarbolas,	Sulphate.	Chloride of Barium.
Spiritus Ætheris Nitrosi,	Deficiency of Nitrite of Ethyl.	Quantitative Analysis.
	Free Acid.	Effervescence with Bicarbonate of Sodium.
	General.	Specific gravity.
Spiritus Ammoniæ,	Empyreumatic substances.	Neutralization with Sulphuric Acid and odor, and Permanganate of Potassium.
	Carbonate.	Effervescence with Acids.
	Sulphate.	Chloride of Barium.
	Chloride.	Nitrate of Silver.
	Calcium.	Oxalate of Ammonium.
	Metals.	Sulphuretted Hydrogen or Sulphide of Ammonium.
Spiritus Frumenti,	Fusel Oil.	Odor on evaporation.
	More than .25 per cent. of solids.	Drying at 100° C.
	Sugar, Glycerin, or spices.	Characters of solids on evaporation.
	Excess of Acid.	Quantitative Analysis.
	Deficiency in alcohol.	Specific gravity.
Spiritus Vini Gallici,	Fusel Oil.	Odor on evaporation.
	Amyl Alcohol.	Sulphuric Acid.
	Methyl Alcohol, Aldehyd, or Oak Tannin.	Solution of Potassa.
	Methyl Alcohol.	Permanganate of Potassium.
Strychnina,	Brucine.	Nitric Acid.
Sulphuris Iodidum,	Mineral matter.	Incineration.
Sulphur Lotum,	Free Acid.	Litmus.
	Arsenious Sulphide.	Wash with Ammonia; evaporate to dryness.
	Arsenious Acid.	Hydrosulphuric Acid.
Sulphur Præcipitatum,	Free Acid.	Litmus.
	Sulphate of Calcium.	Chloride of Barium. Carbonate of Ammonium and Ammonia.
	Alkalies.	Solubility in water.
	Alkaline earths.	Solution in Hydrochloric Acid, and evaporation.
	Arsenious Sulphide.	Ammonic Hydrate.
	Arsenious Acid.	Hydrosulphuric Acid.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Sulphur Sublimatum,	Earthy matter.	Incineration.
Syrupus Acidi Hydriodici,	Free Iodine.	Mucilage of Starch.
	Sulphuric Acid.	Chloride of Barium.
	Hydrochloric Acid.	Nitrate of Silver and Ammonia.
Syrupus Ferri Bromidi,	Free Bromine.	Mucilage of Starch.
Syrupus Ferri Iodidi,	Free Iodine.	Mucilage of Starch.
Tamarindus,	Traces of Copper.	Iron.
Thymol,	Carbolic Acid.	Ferric Chloride to saturated aqueous solution.
Tinctura Ferri Acet.,	Zinc and Copper.	Hydrosulphuric Acid, after removing Iron.
	Fixed Alkalies.	Evaporation and ignition, after removing Iron.
	Ferrous Salt.	Ferricyanide of Potassium.
Tinctura Ferri Chloridi,	Zinc or Copper.	Hydrosulphuric Acid, after removing Iron.
	Fixed Alkalies.	Evaporation and ignition, after removing Iron.
	Nitric Acid.	Sulphuric Acid and Ferrous Sulphate.
	Ferrous Salt.	Ferricyanide of Potassium.
	Oxychloride.	Dilution with water, and boiling.
Veratrina,	Mineral matter.	Incineration.
Vinum Album,	Tannic Acid.	Ferric Chloride.
	Excess or deficiency of Alcohol.	Quantitative Analysis.
	Excess or deficiency of Acid.	Quantitative Analysis.
Vinum Album Fortius,	Excess or deficiency of Alcohol.	Quantitative Analysis.
Vinum Rubrum,	Excess or deficiency of Alcohol.	Quantitative Analysis.
	Excess or deficiency of Acid.	Quantitative Analysis.
	Aniline colors.	Ammonia, Ether; evaporation of ethereal solution in contact with silk.
Zinci Acetas,	Lead or Copper.	Hydrosulphuric Acid.
	Iron, Aluminium, or alkaline earths.	Carbonate of Ammonium in excess.
	Salts of Alkalies or alkaline earths.	Removal of Zinc; evaporation and ignition of filtrate.
Zinci Bromidum,	Lead or Copper.	Hydrosulphuric Acid.
	Iron, Aluminium, or alkaline earths.	Carbonate of Ammonium in excess.

TABLE OF TESTS—*Continued.*

Name of Preparation.	Impurities.	Tests.
Zinci Bromidum,	Alkalies or alkaline earths.	Evaporation and ignition, after removing Zinc.
Zinci Carbonas Præcipitatus,	Lead or Copper.	Hydrosulphuric Acid.
	Iron, Aluminium, or alkaline earths.	Carbonate of Ammonium in excess.
	Salts of Alkalies or alkaline earths.	Evaporation and ignition, after removing Zinc.
Zinci Chloridum,	Basic Chloride.	Alcohol to aqueous solution.
	Lead or Copper.	Hydrosulphuric Acid.
	Iron, etc.	Carbonate of Ammonium in excess.
	Alkalies or alkaline earths.	Evaporation and ignition, after removing Zinc.
Zinci Iodidum, *Zinci Oxidum,*	Same as other Zinc Salts.	
Zinci Phosphidum,	Lead or Copper.	Hydrosulphuric Acid.
Zinci Sulphas,	Chloride.	Nitrate of Silver.
	Same as other Zinc Salts.	
Zinci Valerianas,	Alkalies or alkaline earths.	Evaporation and ignition, after removing Zinc.
	Butyrate of Zinc.	Acetate of Copper.
Zincum,	Arsenic.	Nascent Hydrogen and Nitrate of Silver.
	Lead, Iron, or Copper.	Excess of Ammonia.

Saturation Tables.

Equivalent Weights of Citric Acid, Tartaric Acid, Carbonate of Potassium, Bicarbonate of Potassium, Carbonate of Sodium, Bicarbonate of Sodium, Carbonate of Ammonium, and Carbonate of Magnesium, repeated (in **black**) for 20 parts of each, and incidentally (in Roman) for other proportions. (Exact to two places of decimals.)

Citric Acid ($H_3C_6H_5O_7,H_2O$) ÷ 3 × 2 = 140	**20.00**	18.66	**16.96**	**14.00**	**9.78**	**16.66**	**23.72**	**29.31**
Tartaric Acid $H_2C_4H_4O_6$ = 150	21.43	**20.00**	**18.26**	**15.00**	**10.49**	**17.85**	**25.42**	**31.41**
Carbonate of Potassium K_2CO_3 + 16% Aq. = 164.285	**23.47**	**21.90**	**20.00**	16.43	11.48	19.52	27.87	34.40
Bicarbonate of Potassium 2($KHCO_3$) = 200	**28.57**	**26.66**	24.34	**20.00**	13.98	23.81	33.89	41.90
Carbonate of Sodium $Na_2CO_3,10H_2O$ = 286	**40.08**	**38.13**	34.81	28.60	**20.00**	34.04	48.47	59.98
Bicarbonate of Sodium 2($NaHCO_3$) = 168	**24.00**	**22.40**	20.45	16.80	11.74	**20.00**	28.47	35.18
Carbonate of Ammonium ($N_4H_{16}C_2O_6$) ÷ 2 = 118	**16.85**	**15.73**	14.36	11.80	8.25	14.04	**20.00**	24.71
Carb. of Magnes. (($MgCO_3)_3,Mg2HO,4H_2O$) ÷ 4 = 95.5	**13.64**	**12.73**	11.62	9.55	6.68	11.37	16.18	**20.00**

The amount of acid given in any column will saturate the amount of carbonate in the same column, and *vice versâ*.

The amounts of carbonate in any column are equal to each other in chemical power.

Lemon-juice (sp. gr. 1.039) contains, on an average, 7 per cent. by weight of citric acid.

The same Table in Round Numbers, for Purposes of Prescribing and Dispensing.

(The old names in Latin.)

Citric Acid	**20**	19	**17**	**14**	**10**	**17**	**24**	**30**
Tartaric Acid	22	**20**	**18**	**15**	**11**	**18**	**26**	**32**
Carbonate of Potassium (Potassæ Carbonas)	**24**	**22**	**20**	16	12	20	28	35
Bicarbonate of Potassium (Potassæ Bicarbonas)	**29**	**27**	24	**20**	14	24	34	42
Carbonate of Sodium (cryst.) (Sodæ Carbonas)	**40**	**38**	35	28	**20**	34	49	60
Bicarbonate of Sodium (Sodæ Bicarbonas)	**24**	**22**	20	17	12	**20**	29	36
Carbonate of Ammonium (Ammoniæ Carbonas)	**17**	**16**	14	12	8	14	**20**	25
Carbonate of Magnesium (Magnesiæ Carbonas)	**13**	**13**	11	9	7	11	16	**20**

The Table is read thus: 20 grains of Citric Acid will saturate 29 grains of Bicarbonate of Potassium; 20 grains of Bicarbonate of Sodium will saturate, or be saturated by, 18 grains of Tartaric Acid; 11 grains of Tartaric Acid = 8 grains of Carbonate of Ammonium; 20 grains of Bicarbonate of Sodium are equivalent to, or will do as much work as, 34 grains of Carbonate of Sodium; 14 grains of Citric Acid are as strong as 15 of Tartaric Acid. It is occasionally convenient to double the numbers, halve them, or take some other proportion; also to employ them in weights other than grains.

Lemon-juice contains, on an average, 32½ grains of citric acid in 1 fluidounce, or 4 grains per fluidrachm.

The Proportion by Weight of Absolute or Real Alcohol (C_2H_5HO) *in 100 Parts of Spirits of Different Specific Gravities* (*Fownes*).

Sp. gr. at 60° (15°.5 C.).	Percentage of real alcohol.	Sp. gr. at 60° (15°.5 C.).	Percentage of real alcohol.	Sp. gr. at 60° (15°.5 C.).	Percentage of real alcohol.
0.9991	0.5	0.9511	34	0.8769	68
0.9981	1	0.9490	35	0.8745	69
0.9965	2	0.9470	36	0.8721	70
0.9947	3	0.9452	37	0.8696	71
0.9930	4	0.9434	38	0.8672	72
0.9914	5	0.9416	39	0.8649	73
0.9898	6	0.9396	40	0.8625	74
0.9884	7	0.9376	41	0.8603	75
0.9869	8	0.9356	42	0.8581	76
0.9855	9	0.9335	43	0.8557	77
0.9841	10	0.9314	44	0.8533	78
0.9828	11	0.9292	45	0.8508	79
0.9815	12	0.9270	46	0.8483	80
0.9802	13	0.9249	47	0.8459	81
0.9789	14	0.9228	48	0.8434	82
0.9778	15	0.9206	49	0.8408	83
0.9766	16	0.9184	50	0.8382	84
0.9753	17	0.9160	51	0.8357	85
0.9741	18	0.9135	52	0.8331	86
0.9728	19	0.9113	53	0.8305	87
0.9716	20	0.9090	54	0.8279	88
0.9704	21	0.9069	55	0.8254	89
0.9691	22	0.9047	56	0.8228	90
0.9678	23	0.9025	57	0.8199	91
0.9665	24	0.9001	58	0.8172	92
0.9652	25	0.8979	59	0.8145	93
0.9638	26	0.8956	60	0.8118	94
0.9623	27	0.8932	61	0.8089	95
0.9609	28	0.8908	62	0.8061	96
0.9593	29	0.8886	63	0.8031	97
0.9578	30	0.8863	64	0.8001	98
0.9560	31	0.8840	65	0.7969	99
0.9544	32	0.8816	66	0.7938	100
0.9528	33	0.8793	67		

THE ELEMENTS.

	Symbols and atomic value.	Atomic weight.
Aluminium (Al_2^{VI})	Al^{I}	27
Antimony (Sb^{III}) (119.6, Schneider, Cooke)	Sb^{V}	120
Arsenicum (As^{III})	As^{V}	74.9
Barium	Ba^{II}	136.8
Beryllium (Glucinum)	Be^{II}	9.3
Bismuth (Bi^{III}) (207.5, Dumas)	Bi^{V}	210
Boron (10.9, Berzelius)	B^{III}	11
Bromine (79.75, Stas)	Br^{I}	79.8
Cadmium (111.7, Lenssen)	Cd^{III}	111.8
Cæsium	Cs^{I}	132.7
Calcium (39.9, Erdmann and Marchand)	Ca^{II}	40
Carbon (C^{II})	C^{IV}	12
Cerium (Ce^{III})	Ce^{VI}	138
Chlorine (35.368, Stas)	Cl^{I}	35.4
Chromium (Cr_2^{VI})	Cr^{VI}	52.4
Cobalt (Co^{II})	Co^{VI}	58.6
Copper	Cu^{II}	63.2
Didymium? (138, Mendelejeff)	D^{II}	142.4
Erbium? (171, Mendelejeff)	Eb^{II}	168.9
Fluorine (18.96, Luca, Louget)	F^{I}	19
Gallium	Ga^{IV}	69.8
Germanium (72.3, Boisbandran, 72.75, Winkler).		
Glucinum. See Beryllium.		
Gold (Au^{I}) (196.2, Berzelius)	Au^{III}	196.85
Hydrogen	H^{I}	1
Indium	In^{VI}	113.4
Iodine (126.533, Stas)	I^{I}	126.6
Iridium	Ir^{IV}	192.5
Iron (Fe^{II} & Fe_2^{VI})	Fe^{VI}	55.9
Lanthanum	La^{II}	139.3
Lead ($^{II}Pb^{II}$) (206.4, Stas)	Pb^{IV}	206.5
Lithium	L^{I}	7
Magnesium (23.94, Dumas)	Mg^{II}	24
Manganese Mn^{II} & Mn^{IV})	Mn^{VI}	54.8
Mercury (199 8, Erdmann & Marchand)	Hg^{II}	199.7
Molybdenum	Mo^{VI}	95.9
Nickel (Ni^{II})	Ni^{VI}	58.6
Niobium	Nb^{V}	93.7
Nitrogen (N^{I} & N^{III}) (14.009, Stas)	N^{V}	14
Osmium	Os^{IV}	195
Oxygen (15.96, Stas)	O^{II}	16

	Symbols and atomic value.	Atomic weight.
Palladium	Pd^{IV}	106
Phosphorus (P^{III}) (30.96, Schrötter)	P^{V}	31
Platinum (197.88, Andrews)	Pt^{IV}	194.4
Potassium (39.04, Stas)	K^{I}	39
Rhodium	Rh^{IV}	104.1
Rubidium	Rb^{I}	85.3
Ruthenium	Ru^{IV}	104.2
Scandium.		
Selenium or Selenion	Se^{VI}	78.8
Silicon	Si^{IV}	28.3
Silver (107.66, Stas)	Ag^{I}	107.7
Sodium (22.98, Stas)	Na^{I}	23
Strontium	Sr^{II}	87.4
Sulphur (S^{II} & S^{IV})	S^{VI}	32
Tantalum	Ta^{V}	182
Tellurium	Te^{VI}	125
Thallium	Tl^{III}	203.5
Thorinum or Thorium	Th^{II}	231.4
Tin (Sn^{II})	Sn^{IV}	117.7
Titanium	Ti^{IV}	48
Tungsten	W^{VI}	183.6
Uranium	U^{VI}	239.8
Vanadium	V^{V}	51
Ytterbium	Yb	173
Yttrium	Y^{II}	88.9
Zinc	Zn^{II}	64.9
Zirconium	Zi^{IV}	90.4

The *quantivalence* or atomic value of some elements is, apparently, variable; in the above Table the full coefficients are given in the column of symbols, other common values in parentheses.

Atomic weights are sometimes obscurely termed *equivalents*.

Other elements than the above exist. They are very rare. Some of the rarer so-called elements may not be truly elementary.

Students must expect the figures representing the atomic weights of elements to vary slightly from time to time, in accordance with the advancement of knowledge in the directions of purity of materials and improvements in manipulation, and as regards modes of research and realization of chemical and physical analogies amongst elements.