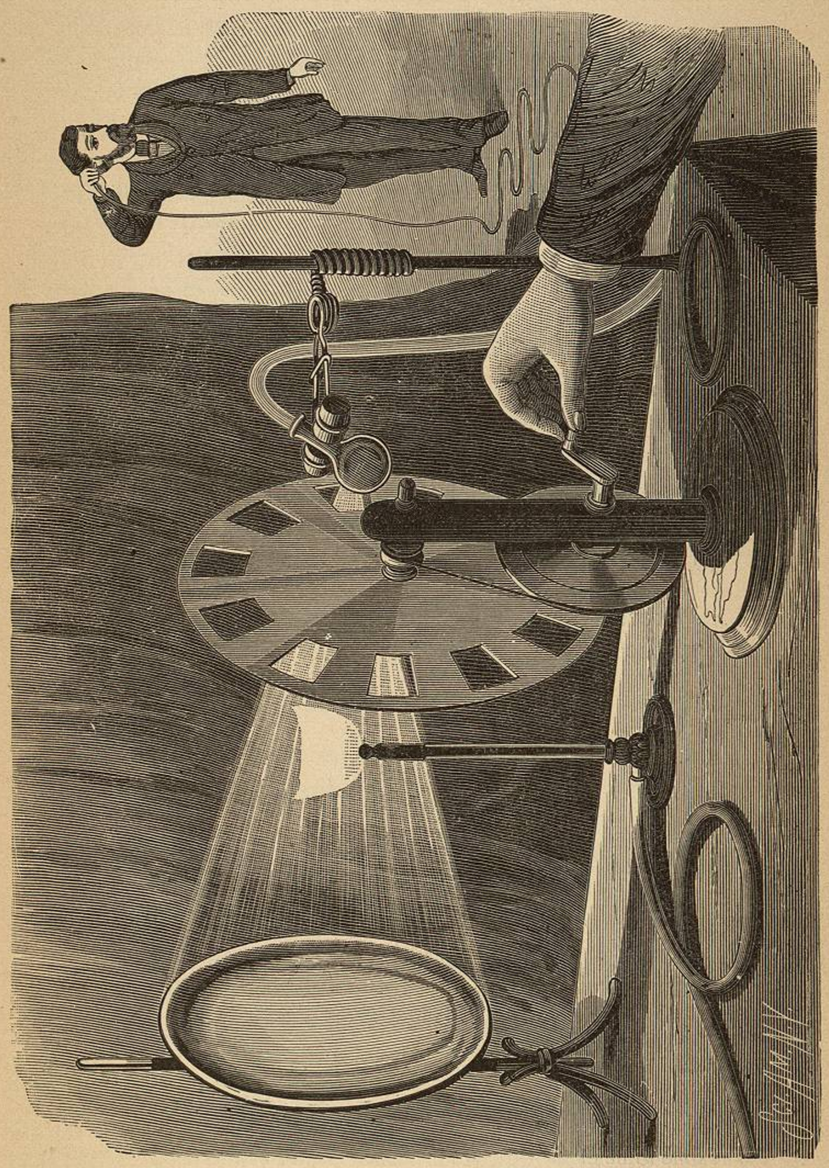
The interesting experiment of Tyndall on radiant heat was suggested to him by Prof. Bell's photophonic experiment, in which musical sounds are obtained by the action of an intermittent beam of light upon a solid body. Referring to this, Prof. Tyndall says:

"From the first I entertained the opinion that these singular sounds were caused by rapid changes of temperature, producing corresponding changes of shape and volume in the bodies impinged upon by the beam. But if this be the case, and if gases and vapors really absorb radiant heat, they ought to produce sounds more intense than those obtained from solids. I pictured every stroke of the beam responded to by a sudden expansion of the absorbent gas, and concluded that when the pulses thus excited followed each other with sufficient rapidity, a musical note must be the result. It seemed plain, moreover, that by this new method many of my previous results might be brought to an independent test. Highly diathermanous bodies, I reasoned, would produce faint sounds, while highly athermanous bodies would produce loud sounds—the strength of the sound being, in a sense, a measure of the absorption. The first experiment, made with a view of testing this idea, was executed in the presence of Mr. Graham Bell, and the result was in exact accordance with what I had foreseen."

The writer has successfully repeated Prof. Tyndall's experiment with the simple apparatus shown in the illustration (Fig. 188). Apparatus already at hand was utilized. A small sized bulbous glass flask, $1\frac{3}{4}$ inches in diameter, was mounted in a test tube holder, and placed behind a rotating pasteboard disk, 12 inches in diameter, having twelve apertures $1\frac{1}{2}$ inches wide and $1\frac{1}{4}$ inches long. Several flasks of the same capacity were provided and filled with the different gases and vapors, and stoppered, to be used at convenience. Near the disk was placed a common gas flame, and into the mouth of the flask was inserted one end of a long rubber tube, the other end being provided with a tapering ear tube, placed in the ear of the listener, whose position was sufficiently remote from the apparatus to avoid any possible disturbance from the revolving disk or the operator. The

Fig. 188.



Apparatus exhibiting the Action of Radiant Heat on Gaseous Matter.

disk being rotated so as to rapidly intercept the thermal and luminous rays of the gas flame and render the rays rapidly intermittent, the effect on the gases and vapors contained by the different bulbs was noted. Dry air produced no sound; moistened, it yielded a distinctly audible tone, corresponding in pitch with the rapidity of the interruptions of the thermal rays.*

Among gases tried, nitrous oxide and illuminating gas yielded the loudest sounds. Among vapors, water and sulphuric ether were most susceptible to the intermittent rays. A candle flame produced distinctly audible sounds in the more sensitive gases, and a hot poker replacing the gas flame yielded the same results.

By using an ordinary concave spun metal mirror, the heat of the flame was satisfactorily projected from a considerable distance. Considering the crudeness of the apparatus and the delicacy of the action which produces the sounds, it appears remarkable that any satisfactory results were obtained, and the experiment shows that any one interested in the finer branches of scientific investigation may often, with the exercise of a little care, enjoy, without material expense, those deeply interesting experiments.

REFLECTION AND CONCENTRATION OF HEAT.

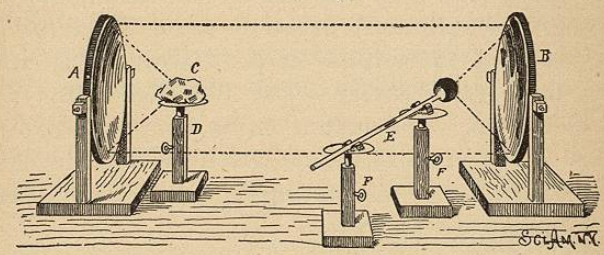
In this experiment the concave mirrors described in a previous chapter are employed in reflecting and concentrating heat.

Instead of placing the watch in the focus of the reflector, B, as in the sound experiment, an air thermometer, E, is supported upon two stands, F F, as shown in Fig. 189, with its bulb in the focus of the reflector. The bulb is smoked over a candle, and when it is nearly cold a drop of water or mercury is introduced into the capillary tube to serve as an index. The candle is removed until the drop in the tube ceases to move. It is then replaced. In a very short time the drop will be pushed outward by the expan-

* The tone to be expected from the gas or vapor when acted on by radiant heat may be determined by blowing through a tube against the apertured portion of the rotating disk.

sion of the air in the bulb. The candle is again removed, and when the drop has returned to the point of starting and ceased moving, a lump, C, of ice is placed on the stand, D,

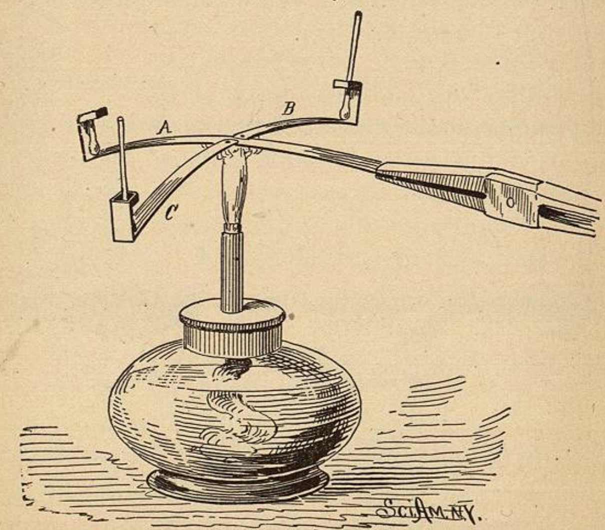
Fig. 189.



Reflection of Heat.

in the focus of the reflector, A. Immediately the air contracts in the thermometer and draws the drop in. Each of the two bodies is radiating, and receiving heat radiated from the other. But the ice radiates less than the bulb; hence the bulb gives out more than it receives, and the fall of temperature is shown by motion of the index.

Fig. 190.



Conduction of Heat.

THE CONDUCTIVITY OF METALS.

The conductivity of metals for heat is admirably shown by the simple device illustrated in Fig. 190. To a strip, A, of iron are attached strips, B C, of brass and copper. The ends of all the strips are bent upward and inward, and the ends of the strips are split and curved to form loops for loosely holding matches, the sulphur ends of which rest upon the strips by their own gravity. The junction of the strips is heated as shown. The match on the copper strip ignites first, that on the brass next, and that upon the iron last, showing that, of the three metals, copper is the best conductor of heat and iron the poorest.

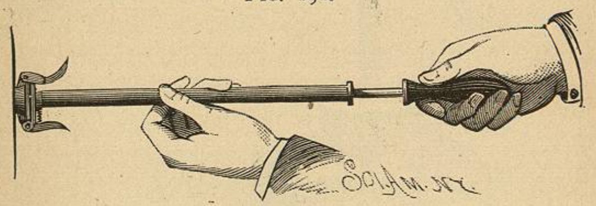
HEAT DUE TO FRICTION.

Every engineer having machinery in charge knows something of this subject. Badly proportioned or poorly lubricated journals often become intensely heated by undue friction. Occasionally a red hot journal is seen. Wherever there is friction there is heat. Often kinetic energy is transformed through friction into heat, which is dissipated by radiation into space, thus causing a loss of energy in a commercial sense, while in a physical sense it still exists, but in another form.

HEAT DUE TO PRESSURE AND COMPRESSION.

Hammering a nail rod until it is red hot and forging a nail without a fire is one of the feats of the blacksmith.

FIG. 191.



Pneumatic Syringe.

The compression of the iron by the blows of the hammer increases its temperature to such a degree as to render this possible. The impact of a bullet on a hard surface generates sufficient heat to melt the lead of which the bullet is formed. Numerous instances might be given of the generation of heat by the impact of solid bodies.

Gases are also heated by compression. By placing some dry tinder or cotton moistened with ether in the pneumatic syringe (pop gun), Fig. 191, and quickly forcing in the piston, so as to strongly compress the air contained in the barrel of the syringe, the temperature of the air will be raised sufficiently to ignite the tinder or cotton.

FIG. 192.

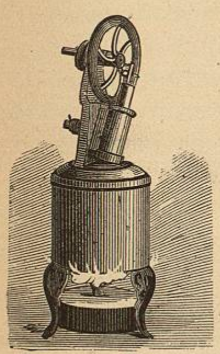
Candle Bomb.

FORCE OF STEAM.

The candle bomb, shown in Fig. 192, exhibits the explosive power of steam. It consists of a small bulb of glass filled with water and sealed. When the bomb is held in a candle flame by means of a wire loop, the water is converted into steam and an explosion occurs.*

The least expensive machine for applying to mechanical work the force exhibited by the candle bomb is the fifty-cent steam engine, shown in Fig. 193. It is a small and simple machine, but it is far more perfect than the steam engines of our forefathers. It will readily make 800 to 1,000 revolutions per minute. It is a wonderfully inexpensive example of the world's greatest motive power. Its construction is so well known that an extended description seems superfluous.

FIG. 193.



Fifty-cent Engine.

The standard which supports the crank shaft also forms the support of the trunnion of the oscillating cylinder. The piston is connected directly with the crank pin projecting from the fly wheel. The face of the cylinder which contacts with the standard forms the valve for admitting steam to the cylinder and releasing it after use. A passage in the standard conveys steam from the boiler to

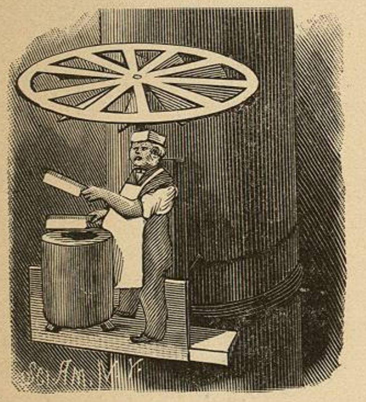
* A guard of some kind should be placed around the bomb to prevent injury to the experimenter.

the steam ports. A spiral spring on the trunnion draws the cylinder against the standard. The cylinder thus arranged is made to serve as a safety valve. A small alcohol lamp is used as a source of heat.

ASCENSIONAL POWER OF HEATED AIR.

The ascensional power of heated air is exhibited by the draught of every chimney. It is shown by the fire balloon and by the upward tendency of every flame. It is the prime factor in the propelling power of one of the most ancient of motors—the windmill; wind being only air rushing forward to take the place of air which is rising because it is rarefied by heat.

Fig. 194.



Hot Air Motor.

The power derived directly from an ascending column of heated air has never been utilized except as a motor for ventilators, for running mechanical toys, and to some extent for operating small mechanical signs.

The toy motor shown in the annexed engraving is too familiar to require description. It is generally placed over a lamp chimney or at the side of a stovepipe, where the rapidly ascending heated air may impinge on the inclined vanes. The air, acting on the vanes according to the well known law of the inclined plane, produces a lateral movement of each vane, and the vanes being restrained at the center of the wheel while free at their outer ends are compelled to move circularly.

HYGROMETRY.

The toy hygroscope serves to show approximately the hygrometric state of the atmosphere. One of the several forms in which it is made is shown in the annexed engraving. A perforated metal tube, projecting from the back of

the figure, contains a short piece of catgut cord, which is fastened in the rear end of the tube by closing the sides of the tube down upon it. The opposite end of the cord projects beyond the front of the figure, and is attached to the arm of the boy. In the hand of the arm thus supported is carried an umbrella. When the air is dry, the catgut cord retains its twist, and the arm holds the umbrella out of the position of use; but when the air becomes moist, the cord swells slightly, and untwists, and in so doing raises the boy's arm and brings the umbrella over his own head and over the head of his companion.

Fig. 195.

Hygroscope.

Another form of the same device consists of a house having two doors and containing two figures—a man with an umbrella and a woman in fair-weather dress; the figures being supported on opposite ends of a bar suspended centrally by a catgut cord. When the cord is untwisted by the action of moisture, the man with the umbrella sallies out; when the cord becomes dry, the man returns indoors and the woman appears.

Fig. 196.

Sensitive Leaf.

These simple, pleasing, and instructive toys illustrate the action of moisture on certain porous bodies, and are of interest, if not of actual use, to the meteorological observer. The action of the sensitive leaf shown in the engraving is also due to expansion by absorption of moisture. The leaf consists of a piece of thin gelatinized paper or gold beater's skin, or even of gelatine, printed in some fantastic design, that of the mermaid being the favorite. When the leaf is laid upon the palm of the hand, the moisture of the hand is absorbed by one side of the leaf, and more in some places than in others, owing to imperfect contact with the hand. The moistened portions rapidly swell, thus warping the leaf, which twists and writhes in every possible direction, as if it were possessed of life. The leaf, being extremely thin, quickly becomes dry, so that the various contortions succeed each other rapidly.

CHEMICAL THERMOSCOPE, HYGROSCOPIC AND LUMINOUS ROSES.

The chemical thermoscope is made by sealing in a tube a solution of chloride of cobalt in dilute alcohol. When the tube is subjected to a temperature of 40° to 50° Fah., the solution becomes pink, and as its temperature is raised to 90° or 100°, it passes through various shades of purple, and finally becomes blue.

FIG. 197.

Chemical Thermoscope.

The same salt applied to an artificial flower, a rose for example, renders it visibly hygroscopic. When the air is humid, the rose is pink; and when the air is warm and dry, the rose will be purple or blue. A solution of the same salt constitutes one of the sympathetic inks.

The luminous rose shown in the same vase with the hygroscopic rose is a beautiful example of the wonderful property of storing light possessed by some bodies. The light-storing property is given the rose by a coating of luminous paint, the basis of which is sulphide of calcium. This rose, if exposed to a strong light during the day, will be luminous throughout the night.

FIG. 198.

Hygroscopic and Luminous Roses.

The exact nature of the change which takes place in the phosphorescent substance while exposed to the light is unknown. It is supposed to be due to

some modifying action of the light, rather than chemical action. It has been ascertained that the phosphorescence takes place *in vacuo* as well as in air. Luminous paint has many practical applications. It is used on buoys, guideposts, gates, etc., to render them visible at night. It is applied to match safes with obvious advantage.